

WORKSHOP 3:
Nomenclature

Name: _____

Section _____

In each box, write the correct formula for the ionic compound which contains the positive ion (cation) listed at the top of the column and the negative ion (anion) listed on the left.

	K^+	Ca^{2+}	Al^{3+}	Zn^{2+}	NH_4^+
Cl^-	(1)				
S^{2-}		(2)			
N^{3-}			(3)		
OH^-				(4)	
SO_4^{2-}					(5)
NO_3^-				(6)	
PO_4^{3-}			(7)		
CO_3^{2-}		(8)			

Name the numbered compounds.

1. _____

2. _____

3. _____

4. _____

5. _____

6. _____

7. _____

8. _____

Name the Compounds

Write Formulas

KI	_____	sodium carbonate	_____
NaNO ₃	_____	water	_____
Mg(OH) ₂	_____	carbonic acid	_____
FeO	_____	lead(II) hydroxide	_____
NH ₄ F	_____	zinc nitrate	_____
NaClO ₃	_____	ammonia	_____
SO ₂	_____	barium phosphide	_____
SnS ₂	_____	aluminum nitride	_____
SiBr ₆	_____	ammonium oxide	_____
FePO ₄	_____	hydrofluoric acid	_____
AgCN	_____	chromic acid	_____
Mg(HCO ₃) ₂	_____	lithium phosphate	_____
HCl(aq)	_____	hydrogen chloride	_____
Ca(ClO) ₂	_____	tin(IV) sulfide	_____
Cu ₂ CO ₃	_____	oxygen	_____
HNO ₃	_____	carbon tetrachloride	_____
N ₂ O ₃	_____	magnesium hydroxide	_____
PbSO ₄	_____	hypobromous acid	_____
AlH ₃	_____	nitrous acid	_____
HC ₂ H ₃ O ₂	_____	copper(II) sulfate	_____
Sn(ClO ₄) ₄	_____	aluminum acetate	_____